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INTRODUCTION

METHODS



Off Label Administration

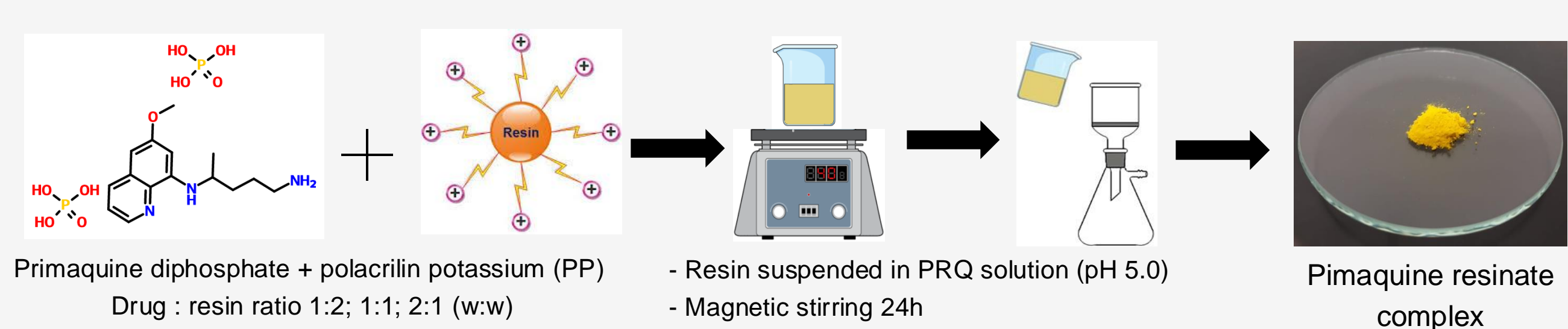
Bitter taste when suspended in water

Ineffective treatment

Non-acceptance of the medicine



Primaquine diphosphate (PRQ) is known for its unpleasant taste, which can lead to non-acceptance of the medicine among malaria patients under five years old, especially when it is given for a radical cure treatment of *P. vivax* and *P. ovale* over 14-day treatment.



Primaquine diphosphate + polacrilin potassium (PP)
Drug : resin ratio 1:2; 1:1; 2:1 (w:w)

- Resin suspended in PRQ solution (pH 5.0)
- Magnetic stirring 24h
- Room temperature

Pimaquine resinate complex

PURPOSE

This study aims to determine the feasibility of primaquine-resin complex formation and evaluate the taste masking efficiency through an adapted dissolution test.

Solid state NMR

- Recycle time of 15 s,
- Spinning rate of 10.0 kHz
- Contact time of 8 ms.

DSC

- 2 – 4mg sample
- Heating rate 10°C/min (30 – 300°C)
- non-hermetic aluminum pan

Dissolution test

- 900mL HCl 0.1N
- pH 1.2; 37 °C
- Paddles at 100 rpm
- UV-Vis measurement at 260nm

Taste Masking

- 500 mL artificial saliva
- pH 6.8 and 37° C
- Paddles 50 rpm
- UV probe measurement at every 5 seconds for 3 minutes.

RESULTS

Drug Load Efficiency

Solid-state NMR

Table 1: PRQ loading efficiency on ion exchange resin (IER) complex.

Drug:resin ratio (w:w)	Drug load efficiency (%)	Conclusion
1:2	95.09 ± 0.005	✗
1:1	91.07 ± 1.03	✓
2:1	25.30 ± 0.02	✗

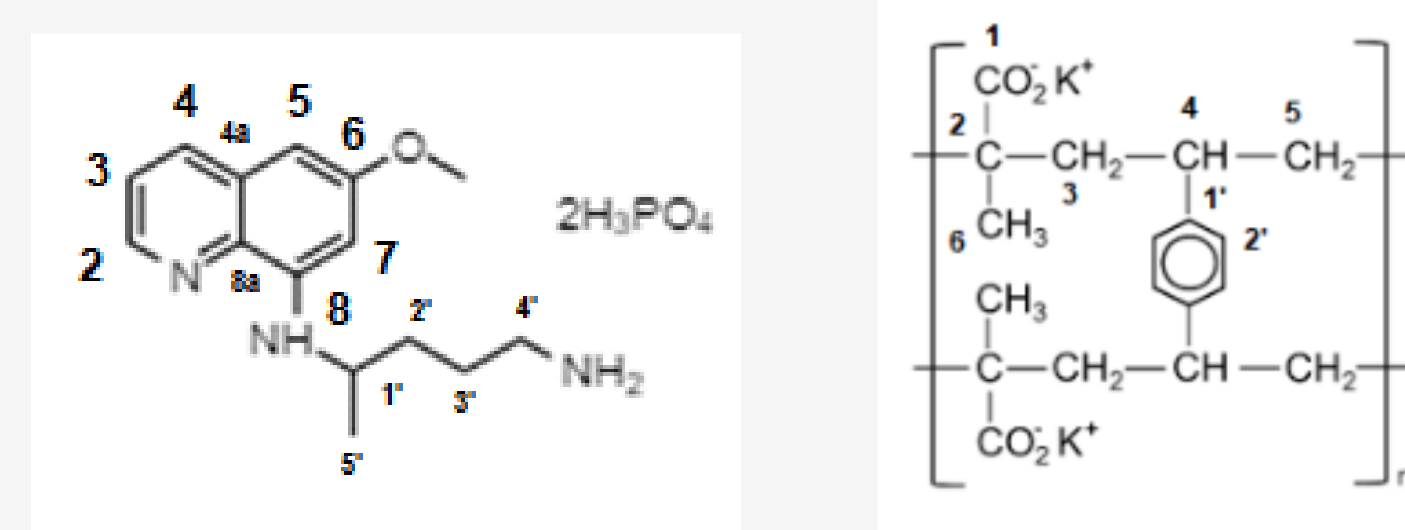


Figure 4: Structural formula of PRQ and PP ion exchange resin.

The drug-to-resin ratio of 1:1 (w:w) demonstrated high drug loading efficiency at 91.07% ± 1.03%. The other ratios resulted in excess PRQ (2:1) or IER (1:2) after 180 minutes.

Table 2: Chemical shifts (δ, ppm) of PRQ, PP resin and PRQ-PP resinate.

CARBON	PRQ	PP resin ^a	PRQ-PP resinate ^a					
	δ, ppm ^a	δ, ppm ^a	δ, ppm					
C6	159.5	162.9	157.9	159.6	160.6	158.1	-	159.0
C8	140.6	141.1	141.8	145.2	146.4	149.7	-	144.0*
C2	138.6	141.4	141.6	144.3	145.2	138.2	-	144.0*
C8a	133.4	126.6	132.3	135.5	135.5	133.0	-	134.7
C4	130.8	147.3	126.2	134.8	135.5	128.1	-	130.2
C4a	123.9	133.8	125.0	130.0	130.2	127.2	-	121.7*
C3	119.6	123.4	123.1	121.8	121.7	122.1	-	121.7*
C7	100.7	106.2	105.5	96.9	96.5	104.5	-	97.7
C5	94.0	97.6	102.1	91.9	96.1	102.1	-	91.6
CH ₂ -O	58.2	58.5	55.4	55.2	55.8	55.4	-	Overlap with PP
C1'	47.6	51.4	51.8	48.1	52.8	51.8	-	Overlap with PP
PC4'	41.2	42.4	43.3	41.7	42.0	41.5	-	39.7**
C2'	33.5	34.7	29.8	34.1	34.6	29.8	-	33.6**
C3'	24.9	26.2	23.4	29.3	31.1	25.6	-	24.6*
C5'	19.6	21.2	19.8	20.5	21.1	19.8	-	Overlap with PP
C1 (CO ₂ K')	-	-	-	-	-	-	186.5; 185.6	184.2; 182.4; 180.6
C1'arom.	-	-	-	-	-	-	148.4	-
C2'arom.	-	-	-	-	-	-	129.6; 127.8	-
C2, C3,	-	-	-	-	-	-	57.4; 46.8	55.5; 45.7
C4, C5	-	-	-	-	-	-	57.4; 46.8	55.5; 45.7
C6	-	-	-	-	-	-	30 - 10	-

^(a)Solid-state ¹³C CPMAS, this work; ^(b)Solution (CDCl₃) ¹³C diprotonated (Clark et al. 1981); ^(c)simulation diprotonated, Toppin 4.0.8.; ^(d)Solution (CDCl₃) free base (Clark et al. 1981); ^(e)simulation, free base, ChemBioDraw Ultra 12.0 (Bronz, 2012); ^(f)simulation, free base, Toppin 4.0.8.
* loss of definition; ** increase in width.

Table 3: Chemical shifts (δ, ppm), and relaxation times T_{CH} and T1pH obtained for samples PRQ, PP and PRQ-PP resinate.

CARBON	PRQ	PP resin	PRQ-PP resinate
	δ, ppm ^a	T _{CH} (μs)	T1pH (ms)
C6	159.5	n.d.	n.d.
C8	140.6	636	265
C2	138.6	1220	336
C8a	133.4	483	246
C4	130.8	n.d.	n.d.
C4a	123.9	n.d.	n.d.
C3	119.6	511	251
C7	100.7	507	246
C5	94.0	497	231
CH ₂ -O	58.2	597	304
C1'	47.6	402	184
C4'	41.2	373	183
C2'	33.5	341	181
C3'	24.9	402	191
C5'	19.6	n.d.	n.d.
C1 (CO ₂ K')	-	-	-
C1'arom.	-	-	-
C2'arom.	-	-	-
C2, C3, C4, C5	-	-	-
C2, C3, C4, C5	-	-	-
C6	-	-	-
C1	186.5	1930	9.7
C1'arom.	148.4	n.d.	n.d.
C2'arom.	129.6	n.d.	n.d.
C2, C3, C4, C5	57.4	99	9.4
C2, C3, C4, C5	57.4	640	9.0
C6	18	355	11.7
C1	182.9	741	8.4
C1'arom.	n.d.	n.d.	n.d.
C2'arom.	n.d.	n.d.	n.d.
C2, C3, C4, C5	n.d.	n.d.	n.d.
C2, C3, C4, C5	n.d.	n.d.	n.d.
C6	266	9.4	9.4

^(a) assignment suggested. * signal not properly defined

Figure 5: ¹³C CPMAS NMR spectra of the samples evaluated. The carbon signals are numbered following the structures presented in Fig. 4.
(*) denotes spinning sidebands.

NMR analysis showed that all the carbon signals could be assigned, including those linked to the IER polacrilin potassium. A strong increase in the width of quinoline moiety, and a decrease in T1pH values confirmed the interaction between PRQ and IER moieties.

CONCLUSION

REFERENCES

- The results showed that complexation with ion exchange resins does not entirely remove primaquine's bitter taste but significantly reduces its perception.
- The complexation of PRQ with IER presented a high drug loading efficiency and the complexation was demonstrated by DSC and by solid - state NMR analysis.
- Proceeding with the following development stages is feasible to create a new and palatable oral formulation of primaquine resinate for children.

ACKNOWLEDGEMENT

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